

**Biochemistry I Fall Term**

**Name:**

pH Titration: Problem Set #1

Given in the problem:

1) Volume of the acid =

2) [NaOH] =

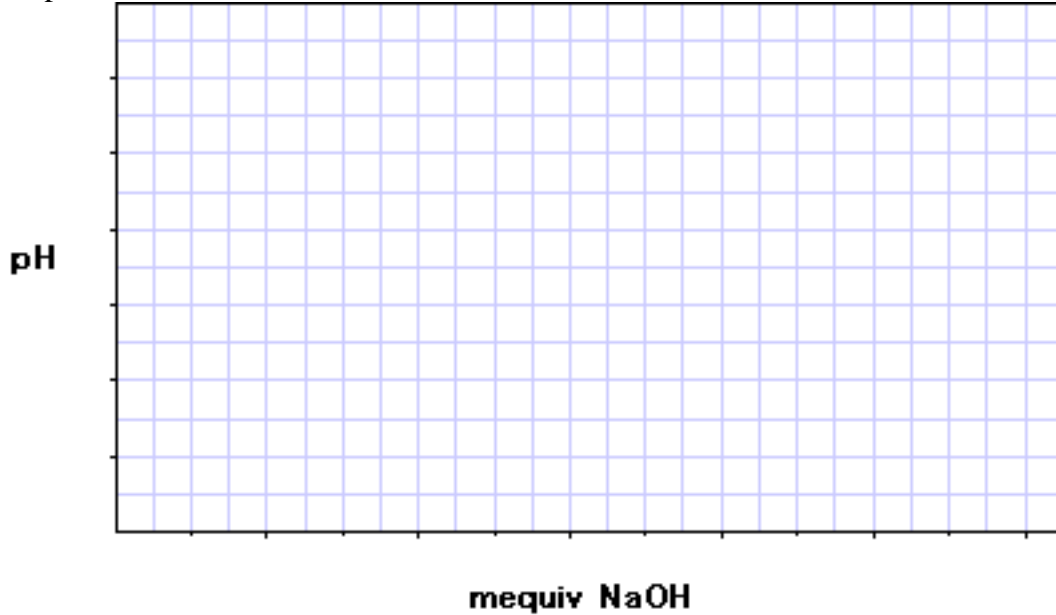
What was found in the experiment:

<b>Vol. NaOH (mL)</b>	<b>pH</b>	<b>mequiv.*</b>
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\*Equivalents of NaOH = Volume x concentration.

mequiv. = mL x 1.0 M (*i.e.* mequiv. =  $10^{-3}$  x equivalents.)

Graph of the calculated results:



The acid concentration = \_\_\_\_\_.

The pK<sub>a</sub> = \_\_\_\_\_.

From Table 2.6 of Campbell, the acid is \_\_\_\_\_.